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## Questions Easily | Ionic Equilibrium

*Solubility Product Constant (K<sub>sp</sub>)*

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$x = 1.33 \times 10^{-5} \text{ M}$ . This is the answer because there is a one-to-one relationship between the Ag<sup>+</sup> dissolved and the AgCl it came from.

So, the molar solubility of AgCl is 1.33

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$\times 10^{-5}$  moles per liter. Calculate the molar solubility (in mol/L) of a saturated solution of the substance.

## ~~SOLUBILITY PROBLEMS~~

Answer. First, treat the solubility equilibrium in units of molarity and then change the concentration to ppb.

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$\text{Cu}_3(\text{AsO}_4)_2(\text{s}) \rightleftharpoons 3\text{Cu}^{2+}(\text{aq}) + 2\text{AsO}_4^{3-}(\text{aq})$   
 $K_{\text{sp}} = [\text{Cu}^{2+}]^3 [\text{AsO}_4^{3-}]^2 = 7.6 \times 10^{-36}$ . Initial 0 0.

Change +3x +2x. Equilibrium 3x 2x.

$7.6 \times 10^{-36} = [3x]^3 [2x]^2 = 108x^5$ .  $x = 3.7 \times 10^{-8} \text{ M}$  = molar solubility of copper(II) arsenate.

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Description Of : Solubility Curve  
Practice Problems Part 2 Answer Key

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Part 2 Answer Key~~

Answers. 1 Solubility Curves . There are charts and tables available that we can use to get an idea of how soluble

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a certain solute is in a certain solvent. We will take a look at two of them in these next two sections. Solubility curves, like the one shown here, tell us what mass of solute will dissolve in 100g (or 100mL; see note

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~~Worksheet 1~~

"Solubility Curve Practice Problems  
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Worksheet. Solubility Curve Practice  
Problems Worksheet 1. Problems  
Worksheet. Solubility Curve

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Worksheet Answer Key. Practice Worksheet.

~~Solubility Curve Practice Problems Worksheet 1 Answer Key ...~~

very small (the solubility is reduced in the presence of a common ion), the term " $0.020 + x$ " is the same as

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"0.020." (You can leave  $x$  in the term and use the quadratic equation but it will not improve the significance of your answer.) :  $1.1 \times 10^{-10} = [x][0.020 + x] = [x][0.020]$   $x = 5.5 \times 10^{-9}$  M Effect of the Common Ion on Solubility

~~Unit 12 Subjects SOLUBILITY~~

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## ~~PRODUCT CALCULATIONS~~

now take the solubility, and multiply it by 325 and divide it by 100 (Rule of three). A few tips for the rest of the problem: 2) Procedure is in reverse order to 1) 3) Subtract the solubilities of...

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Sample Problem #2 If 0.0067g CaCO<sub>3</sub> soluble in 1.0L of water, calculate

K<sub>sp</sub> molar solubility = (0.0067g/L)(1 mol/100g) = 6.7x10<sup>-5</sup> M CaCO<sub>3</sub>(s)

Ca<sup>+2</sup> + CO<sub>3</sub><sup>-2</sup> 6.7x10<sup>-5</sup>M 6.7x10<sup>-5</sup>M

6.7x10<sup>-5</sup>M K<sub>sp</sub> = [Ca<sup>+2</sup>][CO<sub>3</sub><sup>-2</sup>] =

[6.71x10<sup>-5</sup>][6.7x10<sup>-5</sup>] = 4.5x10<sup>-9</sup>

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Sample Problem #3 If 0.017g  $\text{CaF}_2$  soluble in 1.0L of water, calculate  $K_{sp}$

~~$K_{sp}$  Problems — Chemistry~~

SOLUBILITY CURVES Answer the following questions based on the solubility curve below. Which salt is least soluble in water .. 2. How many



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grams of potassium chloride can be dissolved in 200 g of water at 80° C?  
IO 3. At 40° C, how much potassium  
'1 80 70 \_\_\_ nitrate can be dissolved  
^n 30D.g of water?...- O --60-----W- 0  
5© 4. Which salt shows the least  
change 40

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~~SOLUBILITY CURVES PTHS~~

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Solubility Curve Practice Problems

Answer Key SOLUBILITY

PROBLEMS. Here are some practice problems for writing  $K_{sp}$  expressions.

Write the chemical equation showing how the substance dissociates and

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write the  $K_{sp}$  expression: PART 1: 1)  $AlPO_4$  2)  $BaSO_4$  3)  $CdS$  4)  $Cu_3(PO_4)_2$  5)  $CuSCN$  6)  $Hg_2Br_2$  7)  $AgCN$  8)  $Zn_3(AsO_4)_2$  9)  $Mn(IO_3)_2$  10) ...

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? Molar solubility =  $7.07 \times 10^{-7}$  . 5.  $Ag_2$

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$\text{CO}_3^{2-} \rightleftharpoons 2\text{Ag}^+ + \text{CO}_3^{2-}$   $K_{sp} =$   
 $8.1 \times 10^{-12} = 2x \cdot x$   $K_{sp} = [\text{Ag}^+]^2 [\text{CO}_3^{2-}]$   
 $8.1 \times 10^{-12} = 4x^3$   $x^3 =$   
 $2.015 \times 10^{-12}$   $x = 1.3 \times 10^{-4}$  ? Molar  
 solubility is  $1.3 \times 10^{-4}$  . 6.  $\text{AgI} \rightleftharpoons \text{Ag}^+ + \text{I}^-$   
 $\text{NaI} \rightleftharpoons \text{Na}^+ + \text{I}^-$   $x \cdot x \cdot 0.2 = 0.2$

~~Solubility Product Practice Problems~~

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~~Stan's Page~~

Solubility Graph Worksheet Answers

Solubility Product Worksheet -

Answers. 1) What is the concentration of a saturated silver (I) acetate

solution?  $K_{sp}(\text{AgC}_2\text{H}_3\text{O}_2) = 1.94$

$\times 10^{-3}$ . Since  $K_{sp} = [\text{Ag}^+][\text{C}_2\text{H}_3\text{O}_2^-]$ ,

and the concentration of silver ions

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is the same as the concentration of acetate ions, we can set up the

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