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$I H P_{vap} = \exp 14.790 ? T K$ and by the ideal gas law $I Solutions to Chemical and Engineering Thermodynamics, 3e N V RT x 18 \text{ mol} \times 8.314 \times 10^7.5 \text{ (bar ? m}^3 \text{ mol K)}$ $T P = = VV 1 \times 10^{23} ? (10 ? x) 106$ volume taken up by liquid Equating P and P vap we have $x 18 \times 8.314 \times 10^{25} T F 5432.8 x 1 .$

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