

Chem Fax Mole Ratios Answers

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Stoichiometry-Mole Ratios.mov Mole ratio - Intro to stoichiometry Chem Fax Mole Ratios Answers

First we determine the mole ratio (remember unknown species on top), H 2 S: H 2 = 8:8 = Number of moles of given species x mole ratio = number of moles required of unknown. So 3.5 x 1 = 3.5 moles of H 2 S could be produced. II. What would be the mass of this maximum yield? To answer this we need to convert our mole value for H 2 S into a mass value.

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moles Cu reacted = 0.1/63.5 = 1.574 x 10⁻³. 2. mass of silver = 52.1 - 50.1 = 2.00g. moles Ag produced = 2/108 = 8.19 x 10⁻¹. 3. ratio Ag:Cu = 8.19 x 10⁻¹ : 1.574 x 10⁻³ = 819 x 10⁻³ : 1.574 x 10⁻³...

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.... .moles Cu reacted = 0.20g x (1 mol / 63.55g) = 3.15x10⁻³ (2).. mass Ag produced = 50.67g - 50.03g = 0.64g.. . . .mole Ag produced = 0.64g x (1 mol / 107.9g) = 5.93x10⁻³ (3) dividing.. 5.93...

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