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SECTION 12.1 THE ARITHMETIC
OF EQUATIONS (pages 353-358)

This section explains how to
calculate the amount of
reactants required or
product formed in a
nonchemical process. It
teaches you how to interpret
chemical equations in terms

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of interacting moles,
representative particles,
masses, and gas volume at
STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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12.1 THE ARITHMETIC OF
EQUATIONS (pages 353–358)

This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

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This section explains how to
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product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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12.1 THE ARITHMETIC OF
EQUATIONS (pages 353–358)

This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

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Chapter 12 Stoichiometry Guided Reading Answers

Introduce the term stoichiometry in your own words. Stress that stoichiometry allows students to calculate the amounts of chemical substances involved in chemical reactions using information obtained from balanced chemical equations.

12.1 The Arithmetic of Equations 12

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Chapter 3: Stoichiometry -
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3.2 1. True or False? Most
hydrogen atoms have a mass
of 1.008 amu. Justify your
answer. If true, explain why
it is true. If false, what
mass do most hydrogen atoms
have? False, 1.008 amu is
actually hydrogen's average
mass, NO atom of hydrogen
actually has the mass of
1.008 amu. 2.

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Designed for students in Nebo School District, this text covers the Utah State Core Curriculum for chemistry with few additional topics.

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