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Acids Bases and Salts Quiz **ACIDS BASES \u0026**

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**SALTS-FULL CHAPTER || CLASS 10 CBSE CHEMISTRY**

**Acid Bases \u0026amp; salts chapter 5 QUESTION**

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9th std Science 5.Acids, Bases and Salts

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Science Chapter 5: Acids, Bases and Salts -

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Questions pH **???** **???????** | pH value of  
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***answers on acid, base, and salt** || **????,***  
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~~Wonders Acids Bases And Salts Answers~~

Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction.

Examples:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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A and B have pH values of 5 and 8 respectively. Which solution will be basic in nature ...

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Ans:- Salts are classified into four categories on the basis of the strength of acid and base. they are: Salt of strong acid and strong base.e.g NaCl, KCl; Salt of strong acid and weak base. e.g.  $\text{NH}_4\text{Cl}$ ,  $\text{NH}_4\text{I}$ ; Salt of weak acid and strong base. e.g.  $\text{CH}_3\text{COONa}$ ,  $\text{COO}^- \text{Na}^+$  + Salt of weak acid and weak base. e.g.  $\text{CH}_3\text{COO}^- \text{NH}_4^+$  + How is common

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salt found in nature?

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2. Bases react with . acids to produce salts  
and water; water to produce acids and salts;  
salts to produce acids and water; neither  
acids, salts, nor water

~~Acids Bases and Salts Multiple Choice  
Questions and Answers~~

The salts of strong acids and weak bases give  
acidic solution having pH less than 7.

Example,  $\text{NH}_4\text{Cl}$ , Ammonium Chloride will have  
pH less than 7. The salts of weak acids and  
strong bases give basic solution having pH

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more than 7. Example,  $\text{Na}_2\text{CO}_3$ , Sodium Carbonate will have pH more than 7. Question 7.

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Answer A reaction in which an acid and base  
react with each other to give a salt and  
water is termed as neutralization

reaction. For Example: (i)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$   
(ii)  $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$  15.

Give two important uses of washing soda and  
baking soda. Answer Two important uses of  
washing soda are:

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Acids, Bases ...~~

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl. It may be white or light blue or yellow depending upon impurities present in it.

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(A) inorganic acid (B) organic acid (C) soft organic acid (D) strong inorganic acid.

Answer. Answer: (a) 5. When milk of magnesia reacts with acetic acid it produces (A) basic salts (B) acidic salts (C) neutral salt (D)

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complex salt. Answer. Answer: (c) 6. Which of the following phenomena will occur when a small amount of acid is added to water ...

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Turmeric is a natural indicator. On adding

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its paste to acid and base separately, which colours would be observed. (a) Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid and yellow in base (d) Red in acid and blue in base. Answer. Answer: (b) Yellow in acid and red in base

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Acids, Bases ...~~

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concept very well. Acids, Bases and Salts  
Class 7 MCQs Questions with Answers. Choose  
the correct answer: Question 1.

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Acids, Bases ...~~

Acids, Bases and Salts In this first quiz on  
GCSE Chemistry we take a look at one vital  
part of the subject - acids, alkalis, bases  
and salts. In it we find out the difference  
between alkalis and bases, how acids and  
alkalis react with one another, how  
neutralisation can produce salts and the  
relevance of pH values.

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~~GCSE Acids, Bases and Salts | Revise pH  
Values and Alkalis~~

Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

~~Acids, bases and salts. 7th Grade Science  
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Acids, Bases ...~~

Acid + Base  $\rightarrow$  Salt + Water. Question 3:  
Rain becomes acidic due to. carbon dioxide.  
sulphur dioxide. nitrogen dioxide. all of the  
above. Correct Option is : 4. Solution : Rain  
becomes acidic because of carbon dioxide,

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sulphur dioxide and nitrogen dioxide. They dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid ...

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